



## UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS International General Certificate of Secondary Education

| CANDIDATE<br>NAME |                            |                       |
|-------------------|----------------------------|-----------------------|
| CENTRE<br>NUMBER  |                            | CANDIDATE<br>NUMBER   |
| CHEMISTRY         |                            | 0620/02               |
| Paper 2           |                            | October/November 2008 |
|                   |                            | 1 hour 15 minutes     |
| Candidates ans    | wer on the Question Paper. |                       |

## **READ THESE INSTRUCTIONS FIRST**

No Additional Materials are required.

Write your Centre number, candidate number and name in the spaces at the top of this page.

Write in dark blue or black pen.

You may need to use a pencil for any diagrams, graphs or rough working.

Do not use staples, paper clips, highlighters, glue or correction fluid.

DO NOT WRITE IN ANY BARCODES.

Answer all questions.

A copy of the periodic table is printed on page 16.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [ ] at the end of each question or part question.

| For Exam | iner's Use |
|----------|------------|
| 1        |            |
| 2        |            |
| 3        |            |
| 4        |            |
| 5        |            |
| 6        |            |
| 7        |            |
| Total    |            |

This document consists of **16** printed pages.



1 (a) The table gives some information about five elements, A, B, C, D and E. Complete the table by writing either metal or non-metal in the last column.

For Examiner's Use

| element   | properties   | metal or non-metal |
|---|--|--------------------|
| A shiny solid which conducts electricity        |  |                    |
| В   | reddish brown liquid with a low boiling point                      |                    |
| С   | a form of carbon which is black in colour and conducts electricity |                    |
| D   | white solid which is an insulator and has a high melting point     |                    |
| E dull yellow solid which does not conduct heat |  |                    |

| ,, | <b>┐</b> I |
|----|------------|
| I١ | וע         |
| L  |            |
|    |            |

| (b) Describe how metallic character changes across a P |
|--|
|--|

[1]

(c) Sodium is in Group I of the Periodic Table.

(i) Draw a diagram to show the full electronic structure of sodium.

[1]

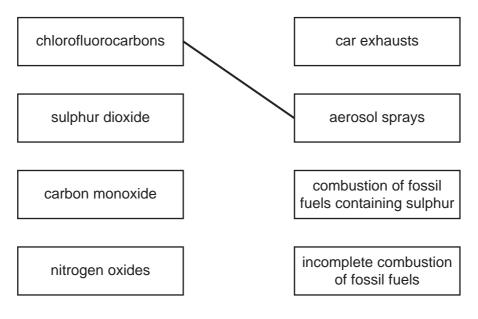
(ii) Complete the equation to show what happens when a sodium atom forms a sodium ion.

| (d) Complete these sentences about the properties of the Group I elements using words<br>from the list. |                          |                    |                      | - 1   | For<br>Examiner's<br>Use |  |
|---|--------------------------|--------------------|----------------------|-------|--------------------------|--|
| acidic  | basic                    | decre              | ase                  | hard  |                          |  |
|   | increase                 | lithium            | potassium            | soft  |                          |  |
|   |                          |                    |                      |       |                          |  |
| Γhe Group   | o I elements are relativ | ely                | metals which         |       | in                       |  |
| eactivity o   | going down the Group     | Sodium reacts more | violently with water | than  |                          |  |
| The Group   | o I metals all form      |                    | oxides.              |       | [4]                      |  |
|   |                          |                    |                      | [Tota | ıl: 12]                  |  |

**2 (a)** Match up the atmospheric pollutants on the left with their main source on the right. The first one has been done for you.

For Examiner's Use

[3]



**(b)** One stage in the manufacture of sulphuric acid involves the oxidation of sulphur dioxide by oxygen in the air to form sulphur trioxide.

$$2SO_2 + O_2 \longrightarrow 2SO_3$$

(i) Explain how this reaction shows that sulphur dioxide is oxidized.

[1]

(ii) What is the percentage of oxygen in clean air? [1]

(iii) Sulphuric acid is used to make the fertiliser ammonium sulphate.

ammonia + sulphuric acid → ammonium sulphate

What type of reaction is this?

[1]

| (iv) | Why do farmers need to use fertilisers?   |     | For<br>Examiner'<br>Use |
|------|---|-----|-------------------------|
|      |   | [2] |                         |
| (v)  | Another fertiliser can be made by the reaction of ammonia with nitric acid. State the chemical name of this fertiliser. |     |                         |
|      |   | [1] |                         |
|      | [Total:   | 91  |                         |

| 3 | Calcium carbonate, | CaCO <sub>3</sub> , is the | raw material | used in the | manufacture o | f lime, | CaC |
|---|--------------------|----------------------------|--------------|-------------|---------------|---------|-----|
|---|--------------------|----------------------------|--------------|-------------|---------------|---------|-----|

For Examiner's Use

(a) (i) Describe how lime is manufactured from calcium carbonate.

| 11! |
|-----|
|     |

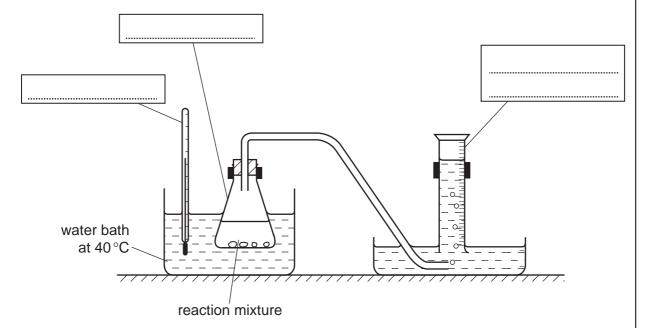
(ii) Write a symbol equation for this reaction.

[1]

(iii) State one large scale use of lime.

[1]

- **(b)** A student investigated the speed of reaction of calcium carbonate with hydrochloric acid using the apparatus shown below.
  - (i) Complete the labelling of the apparatus by filling in the three boxes. [3]



(ii) The equation for the reaction is

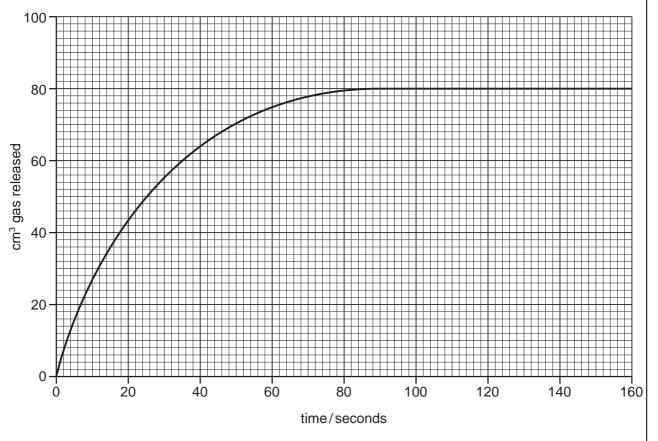
$$CaCO_3 + 2HCl \longrightarrow CaCl_2 + CO_2 + H_2O$$

Write the word equation for this reaction.

[2]

(iii) The student carried out the reaction at 40°C using large pieces of calcium carbonate. The results of the experiment are shown below.

For Examiner's Use



| [1    | Ι. |
|-------|----|
| <br>L | •  |

- (iv) The student repeated the experiment using the same mass of powdered calcium carbonate. All other conditions were kept the same. On the grid above, sketch the graph for the reaction with calcium carbonate powder. [2]
- (v) How does the speed of reaction change when

the concentration of hydrochloric acid is decreased,

the temperature is increased? [2]

[Total: 13]

For Examiner's Use

[1]

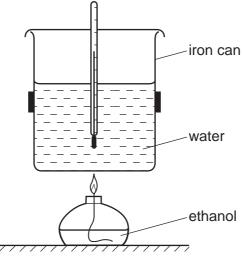
|   |      |      | •  |     |  |
|---|------|------|--|-----|--|
| 4 | Iron | is e | extracted from its ore in a blast furnace.   |     |  |
|   | (a)  | Sta  | te the name of the ore from which iron is extracted.   |     |  |
|   |      |      |  | [1] |  |
|   |      |      |  |     |  |
|   | (b)  | The  | e diagram shows a blast furnace.   |     |  |
|   |      |      | coke + limestone + iron ore  B  air in   |     |  |
|   |      | (i)  | Which <b>one</b> of the raw materials is added to the blast furnace to help remove the impurities from the iron ore?   |     |  |
|   |      |      |  | [1] |  |
|   |      | (ii) | The impurities are removed as a slag. Which letter on the diagram shows the sla  | g?  |  |
|   |      |      |  | [1] |  |
|   | (c)  | Car  | bon monoxide is formed in the blast furnace by reaction of coke with oxygen.   |     |  |
|   | ` ,  |      | Complete the equation for this reaction.   |     |  |
|   |      | \-/  | and the state of t |     |  |
|   |      |      | C +CO  | [2] |  |
|   |      | (ii) | State the adverse affect of carbon monoxide on human health.   |     |  |
|   |      |      |  |     |  |

| (d) | In the hottest regions of the blast furnace the following reaction takes place.  |          | For             |
|-----|--|----------|-----------------|
|     | $Fe_2O_3 + 3C \longrightarrow 2Fe + 3CO$   | <b>I</b> | aminer's<br>Use |
|     | Which two of these sentences correctly describe this reaction? Tick <b>two</b> boxes.  |          |                 |
|     | The iron oxide gets reduced.   |          |                 |
|     | The reaction is a thermal decomposition.   |          |                 |
|     | The carbon gets oxidised.  |          |                 |
|     | The carbon gets reduced.   |          |                 |
|     | Carbon neutralises the iron oxide.   | [1]      |                 |
| (e) | Aluminium cannot be extracted from aluminium oxide in a blast furnace.  Explain why aluminium cannot be extracted in this way. |          |                 |
|     |  | [2]      |                 |
| (f) | (i) State the name of the method used to extract aluminium from its oxide ore.   |          |                 |
|     |  | [1]      |                 |
|     | (ii) State one use of aluminium.   |          |                 |
|     |  | [1]      |                 |
|     | [Total   | : 11]    |                 |
|     |  |          |                 |

5 The apparatus shown below can be used to measure the energy released when a liquid fuel is burnt. The amount of energy released is calculated from the increase in temperature of a known amount of water.

For Examiner's Use

[2]



| (a) | (i)  | (i) Explain how this experiment shows that the burning of ethanol is an exotherm reaction.            |     |  |  |
|-----|------|---|-----|--|--|
|     |      |   | [1] |  |  |
|     | (ii) | Complete the word equation for the complete combustion of ethanol.                                    |     |  |  |
|     |      | ethanol + oxygen → +  | [2] |  |  |
| (b) |      | anol is a fuel containing carbon.<br>te the names of two other commonly used fuels containing carbon. |     |  |  |
|     |      | andand  | [2] |  |  |
| (c) | Giv  | e the formula of the functional group present in ethanol.   | [1] |  |  |
|     |      |   |     |  |  |
| (d) | The  | e can contains water. Describe a chemical test for water.   |     |  |  |

......

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result

| (e) | The   | e iron can used in this experiment rusts easily.  | For               |
|-----|-------|---|-------------------|
|     | (i)   | Describe a method which can be used to prevent iron from rusting.                               | Examiner's<br>Use |
|     |       | [1  | ]                 |
|     | (ii)  | Rust contains hydrated iron(III) oxide. What do you understand by the term <i>hydrated</i> ?    |                   |
|     |       | [1  | ]                 |
|     | (iii) | Iron is a transition metal. State <b>two</b> properties which are typical of transition metals. |                   |
|     |       | [2  | 1                 |
|     |       | [Total: 12  | _                 |

6 The compound shown below is the first member of the alkane homologous series.

| For        |
|------------|
| Examiner's |
| Use        |

| (a) | State two | characteristics | of a | homologous | s series |
|-----|-----------|-----------------|------|------------|----------|

| <br> |
|------|
| [2]  |

| (b) | Name and draw the stru | cture of the next n | nember of the alka | ane homologous series |
|-----|------------------------|---------------------|--------------------|-----------------------|

| name |  |
|------|--|
|      |  |

structure

[2]

(c) Complete the table to show the structure and uses of some organic compounds.

| name of compound | molecular formula                            | structure<br>(showing all atoms and bonds) | use           |
|------------------|--|--|---------------|
| ethene           | C₂H₄   |  |               |
| ethanoic acid    | C <sub>2</sub> H <sub>4</sub> O <sub>2</sub> |  | making esters |
| dibromoethane    |  | Br Br<br>   <br>H—C—C—H<br>   <br>H H      |               |
|                  | CH₄  | H<br> <br>H—C—H<br> <br>H                  |               |

[6]

(d) Calculate the relative molecular mass of dibromoethane.

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[1]

[Total: 11]

7 The diagram shows the structures of calcium chloride, calcium and chlorine.

(ii) At room temperature, calcium is a solid but chlorine is a gas.

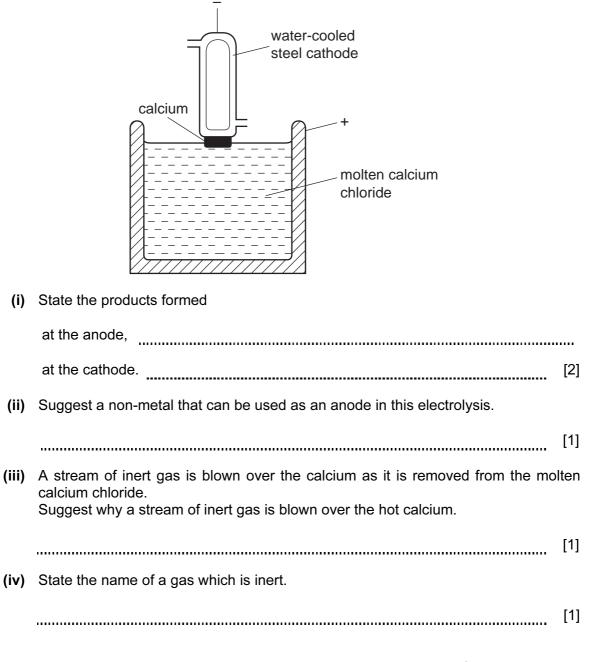
For Examiner's Use

[2]

| $\begin{array}{c c} Cl^{-} & Cl^{-} & Cl^{-} & Cl^{-} \\ \hline Ca^{2+} & Ca^{2+} \\ \hline Cl^{-} & Cl^{-} & Cl^{-} & Cl^{-} \\ \hline Ca^{2+} & Ca^{2+} \\ \hline \end{array}$ | Ca) Ca) Ca)<br>(Ca) Ca) Ca)<br>(Ca) Ca) Ca) | $ \begin{array}{c} Cl \\ Cl \\$ |
|--|---|---|
| calcium chloride   | calcium                                     | chlorine  |
| (a) Use ideas about structure an   | d bonding to explain the follo              | wing:   |
| (i) Calcium chloride conduc  | ts electricity when molten but              | t not when solid.   |
|  |   |   |
|  |   | [2]   |

(b) Calcium is manufactured by the electrolysis of molten calcium chloride.





**(c)** Aqueous sodium hydroxide or aqueous ammonia can be used to test for calcium ions in solution.

Describe the results of these tests

| with aqueous sodium hydro | xide, |     |
|---------------------------|-------|-----|
|                           |       | [2] |
| with aqueous ammonia.     |       |     |
|                           |       | [1] |
|                           |       |     |

[Total: 12]

The volume of one mole of any gas is 24 dm<sup>3</sup> at room temperature and pressure (r.t.p.).

DATA SHEET
The Periodic Table of the Elements

|       | 0        | 4 <b>He</b> ium           | 20<br>Neon<br>10<br>Ar<br>Argan                | 84<br>Krypton<br>36                | 131<br><b>Xe</b><br>Xenon<br>54     | <b>Rn</b><br>Radon<br>86            |                                | Lu<br>Lutetium<br>71                                | <b>Lr</b><br>Lawrencium<br>103   |                                   |
|-------|----------|---------------------------|--|------------------------------------|-------------------------------------|-------------------------------------|--------------------------------|---|--|-----------------------------------|
|       | IIA      |                           | 19 Fluorine 9 35.5 <b>C1</b> Chlorine          | 80<br><b>Br</b><br>Bromine<br>35   | 127 <b>I</b> lodine                 | At<br>Astatine<br>85                |                                | 173 <b>Yb</b> Ytterbium 70                          | No Nobelium 102  |                                   |
|       | IN       |                           | 16<br>Oxygen<br>8<br>32<br><b>S</b><br>Sulphur | Selenium 34                        | 128 <b>Te</b> Tellurium             | Polonium<br>84                      |                                | 169<br><b>Tm</b><br>Thulium<br>69                   | Md<br>Mendelevium<br>101   |                                   |
|       | >        |                           | Nitrogen 7 31 <b>Ph</b> Phosphorus 15          | AS<br>Arsenic                      | 122<br><b>Sb</b><br>Antimony<br>51  | 209 <b>Bi</b> Bismuth 83            |                                | 167<br><b>Er</b><br>Erbium<br>68                    | Fm<br>Fermium  |                                   |
|       | <u>\</u> |                           | 12 Carbon 6 28 Silicon 14                      | 73<br><b>Ge</b><br>Germanium<br>32 | 119<br><b>Sn</b><br>Tin             | 207<br><b>Pb</b><br>Lead            |                                | 165<br><b>Ho</b><br>Holmium<br>67                   | <b>ES</b><br>Einsteinium<br>99   |                                   |
|       | =        |                           | 11 <b>B</b> Boron  27 <b>A1</b> Atuminium  13  | 70 <b>Ga</b> Gallium 31            | 115<br><b>In</b><br>Indium          | 204<br><b>T t</b><br>Thallium<br>81 |                                | 162<br><b>Dy</b><br>Dysprosium<br>66                | Cf<br>Californium<br>98  |                                   |
|       |          |                           |  | 65 <b>Zn</b> Zinc 30               | 112<br><b>Cd</b><br>Cadmium<br>48   | 201<br><b>Hg</b><br>Mercuny<br>80   |                                | 159<br><b>Tb</b><br>Terbium<br>65                   | Bk<br>Berkelium<br>97  |                                   |
|       |          |                           |  | 64<br>Copper                       | 108<br><b>Ag</b><br>Silver<br>47    | 197<br><b>Au</b><br>Gold            |                                | 157<br><b>Gd</b><br>Gadolinium<br>64                | Cm<br>Curium   |                                   |
| Group |          |                           |  | 59<br>Nickel                       | 106 Pd Palladium 46                 | 195 <b>Pt</b> Platinum 78           |                                | 152<br><b>Eu</b><br>Europium<br>63                  | Am<br>Americium<br>95  |                                   |
| Gre   |          |                           |  | 59<br><b>Co</b><br>Cobalt<br>27    | 103<br><b>Rh</b><br>Rhodium<br>45   | 192 <b>Ir</b><br>Iridium            |                                | Sm<br>Samarium<br>62                                | Pu<br>Plutonium<br>94  |                                   |
|       |          | 1<br><b>T</b><br>Hydrogen |  | 56<br>Iran                         | Ruthenium                           | 190<br><b>Os</b><br>Osmium<br>76    |                                | Pm<br>Promethium<br>61                              | Neptunium 93   |                                   |
|       |          |                           |  | Mn<br>Manganese<br>25              | Tc<br>Technetium<br>43              | 186<br><b>Re</b><br>Rhenium<br>75   |                                | 144 <b>Ne</b> Neodymium 60                          | 238<br><b>U</b><br>Uranium<br>92                                       |                                   |
|       |          |                           |  | 52<br><b>Cr</b><br>Chromium<br>24  | 96<br><b>Mo</b><br>Molybdenum<br>42 | 184<br><b>W</b><br>Tungsten<br>74   |                                | Pr<br>Praseodymium<br>59                            | Pa<br>Protactinium<br>91   |                                   |
|       |          |                           |  |                                    | 51<br>V<br>Vanadium<br>23           | 93<br><b>Nb</b><br>Niobium<br>41    | 181 <b>Ta</b> Tantalum 73      |   | 140 <b>Ce</b> Cerium   | 232<br><b>Th</b><br>Thorium<br>90 |
|       |          |                           |  | 48 <b>Ti</b> Tatanium              | 2 <b>r</b><br>Zirconium<br>40       | 178 <b>Hf</b> Hafnium * 72          |                                | 1   | nic mass<br>Ibol<br>nic) number  |                                   |
|       |          |                           |  | Scandium 21                        | 89 <b>×</b>                         | 139 <b>La</b> Lanthanum 57 *        | 227<br><b>AC</b><br>Actinium † | d series<br>series                                  | a = relative atomic mass  X = atomic symbol b = proton (atomic) number |                                   |
|       | =        |                           | Be Beryllium 4 24 Mg Magnesium 12              | 40 <b>Ca</b> Calcium               | Strontium                           | 137<br><b>Ba</b><br>Barium<br>56    | 226<br><b>Ra</b><br>Radium     | *58-71 Lanthanoid series<br>190-103 Actinoid series | <i>a</i> <b>×</b> <i>a</i>   |                                   |
|       | _        |                           | Lithium 3 Lithium 3 23 Na Sodium 11            | 39 <b>K</b> Potassium              | Rb Rubidium                         | Caesium 55                          | Francium<br>87                 | *58-71 L  | Key  |                                   |

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